

JASWANT MODERN SR SEC SCHOOL X WORKSHEET

SET-A CHEMISTRY

Chapter-1

- Q1. Define Chemical Reactions. Give two examples.
- Q2. Write the chemical formula of
(a) Zinc Sulphate (b) Glucose
(c) Calcium Carbonate (d) Formic Acid
- Q3. What is oxidizing agents?
- Q4. Differentiate between Concentrated and Dilute acid
- Q5. Give two definitions of Oxidation with example?
- Q6. Balance the following chemical reaction step wise:
(a) $\text{Fe} + \text{H}_2\text{O} \rightarrow \text{Fe}_3\text{O}_4 + \text{H}_2$
(b) $\text{CO} + \text{H}_2 \rightarrow \text{CH}_3\text{OH}$
- Q7. Write the chemical reaction involved when acid react with metal
- Q8. Write the balanced chemical equations for the following reactions.
(a) $\text{Ca}(\text{OH})_2 + \text{CO}_2 \rightarrow \text{CaCO}_3 + \text{Water}$
(b) Barium chloride + Potassium sulphate \rightarrow Barium sulphate + Potassium chloride
- Q9. Explain the Redox Reaction with the help of example?
- Q10. Why does the colour of copper sulphate solution change when an iron nail is dipped in it? Also write down the chemical reaction involved?
- Q11. How many types of chemical reactions are there? Give one example of each type?
- Q12. What is rancidity? Give methods to prevent it?
- Q13. What are the characteristics of chemical reaction?
- Q14. Write a balanced chemical equation with state symbols for the following reactions.
(i) Solutions of barium chloride and sodium sulphate in water react to give insoluble barium sulphate and the solution of sodium chloride.
(ii) Sodium hydroxide solution (in water) reacts with hydrochloric acid solution (in water) to produce sodium chloride solution and water.

Chapter-2

- Q15. State the colour of phenolphthalein in soap solution
- Q16. Why silver Chloride turns grey in sunlight?
- Q17. Identify the substances that are oxidised and the substances that are reduced in the following reactions.
(i) $4\text{Na}(\text{s}) + \text{O}_2(\text{g}) \rightarrow 2\text{Na}_2\text{O}(\text{s})$
(ii) $\text{CuO}(\text{s}) + \text{H}_2(\text{g}) \rightarrow \text{Cu}(\text{s}) + \text{H}_2\text{O}(\text{l})$
- Q18. Lime water react with chlorine to form
- Q19. What does one mean by exothermic and endothermic reactions? Give examples.
- Q20. Why acids do not show acidic behavior in absence of water?